

- Heat is evolved during:
 - Combination Reaction
 - Combustion Reaction
 - Endothermic Reaction
 - Displacement Reaction
- Dissolving sugar in water is an example of
 - Physical Change
 - Redox Reaction
 - Chemical change
 - None
- Which among the following is not a physical change?
 - Melting of solids to liquids
 - Vaporisation of liquids to gases
 - Liquefaction of gases to liquids
 - Decay of matter
- Chemical changes are _____.
 - temporary, reversible and a new substance is produced
 - always accompanied by exchange of light
 - permanent, irreversible and a new substance is produced
 - never accompanied by exchange of light and heat energy
- Which of the following information is conveyed by a chemical reaction?
 - The colour changes taking place
 - The structure of the reactants and products
 - The absorption of energy only
 - The masses of the reactants and products involved in the reaction
- Chemically rust is
 - Hydrated ferric oxide
 - Hydrated ferrous oxide
 - only ferrous oxide
 - only ferric oxide
- The formula for rust is _____.
 - CuO
 - $\text{Fe}_2\text{O}_3 \cdot x\text{H}_2\text{O}$
 - Al_2O_3
 - AgS
- The reaction $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$ represents
 - Combustion
 - Combination
 - Reduction
 - Oxidation
- Oxidation is a process which involves:
 - Addition of oxygen
 - Removal of hydrogen
 - Both
 - None
- Some crystals of copper sulphate were dissolved in water. The colour of the solution obtained would be:
 - green
 - red
 - blue
 - brown

11. When dilute hydrochloric acid is added to zinc pieces taken in a test tube:
- no change takes place
 - the colour of the solution becomes yellow
 - a pungent smelling gas gets liberated
 - small bubbles of hydrogen gas appear on the surface of zinc pieces.
12. The main cause of rancidity in foods is
- Bacteria
 - Proteins
 - Antioxidants
 - Oxidation of the fatty acid molecule
13. Which of these metals do not corrode?
- Lead
 - Copper
 - Platinum
 - Silver
14. The products of a burning candle are:
- ash and water vapour
 - CO₂ and water vapour
 - wax and water vapour
 - only melted wax
15. CuSO₄ reacts with Zn to form a zinc sulphate and copper. In this reaction, copper act as a:
- Oxidizing agent
 - Reducing agent
 - Dehydrating agent
 - Bleaching agent
16. When KOH is treated with HCl, the reaction is called
- Displacement reaction
 - Double displacement reaction
 - Combination reaction
 - Neutralization reaction
17. The arrangement of metal in the decreasing order of reactivity is
- K>Cu>Ni>Pb
 - Pb>Ni>Cu>K
 - K>Ni>Pb>Cu
 - Ni>K>Cu>Pb
18. $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$ is an example of _____.
- neutralization reaction
 - redox reaction
 - double displacement reaction
 - decomposition reaction
19. $\text{BaCl}_2 + \text{ZnSO}_4 \rightarrow \text{ZnCl}_2 + \text{BaSO}_4$ In the reaction, the white precipitate seen is due to:
- ZnCl₂
 - BaSO₄
 - BaCl₂
 - ZnSO₄
20. Breaking of lead bromide into lead and bromine is an example of:
- decomposition reaction
 - synthesis reaction
 - displacement reaction
 - neutralisation reaction